Association of Chemistry Teachers
NATIONAL STANDARD EXAMINATION IN CHEMISTRY 2011 -2012

Date of Examination 27th November 2011

Time 12.30 to 14.30 Hrs

Instructions to Candidates

01. On the answer sheet, fill up all the entries carefully in the space provided, **ONLY IN BLOCK CAPITALS.** Use only **BLUE or BLACK BALL PEN** for making entries and marking answers. **Incomplete / incorrect / carelessly filled information may disqualify your candidature.**

02. The question paper contains 80 multiple-choice questions. Each question has 4 options, out of which only one is correct. Choose the correct answer and mark a **cross** in the corresponding box on the answer sheet as shown below :

Q.	a	b	c	d
22			X	

03. A correct answer carries 3 marks and 1 mark will be deducted for each wrong answer.

04. All rough work may be done on the blank sheet provided at the end of the question paper.

05. ***PLEASE DO NOT MAKE ANY MARK OTHER THAN (X) IN THE SPACE PROVIDED ON THE ANSWER SHEET.***

    Answer sheets are evaluated with the help of a machine. Due to this, **CHANGE OF ENTRY IS NOT ALLOWED.**

06. **Scratching or overwriting may result in wrong score.**
    DO NOT WRITE ANYTHING ON BACK OF ANSWER SHEET.

07. Use of a nonprogrammable calculator is allowed.

08. Periodic table and log table are provided at the end of this question paper.

09. The answers / solutions to this question paper will be available on our website - **www.iapt.org.in** by 3 rd December 2011.

Certificates & Awards

i) Certficates to top 10% students of each centre.

ii) Merit certificates to statewise Top 1% students.

iii) Merit certificate and a prize in the form of a book to Nationwise Top 1% students.

10. Result sheets and the "centre top 10%" certificates of NSEC are dispatched to the professor in charge of the centre. Thus you will get your marks from the Professor in charge of your centre by January 2012 end.

11. TOP 300 (or so) students are called for the next examination - Indian National Chemistry Olympiad (INChO). Individual letters are sent to these students ONLY.

12. Gold medals will be awarded to TOP 35 students in this entire process.

13. No querries will be entertained in this regard.

NSEC-2011

(1) The number of water molecules present in 0.20g sample of $CuSO_4.5H_2O$ (Molar mass = 249.7) is

(A) 1.2×10^{21} (B) 2.14×10^{21} (C) 2.14×10^{22} (D) 1.2×10^{23}

(2) The group that has the species correctly listed in the order of decreasing radius is

(A) Cu^{2+}, Cu^{+}, Cu (B) V, V^{2+}, V^{3+} (C) F^{-}, Br^{-}, I^{-} (D) B, Be, Li

(3) The number of isomers of dibromobiphenyl (Biphenyl - C_6H_5-C_6H_5) is

(A) 8 (B) 10 (C) 12 (D) 14

(4) The enthalpies of decomposition of methane $(CH_4)(g)$ and ethane $(C_2H_6)(g)$ are 400 and 670 $kJ.mol^{-1}$, respectively. The ΔH_{C-C} in $kJmol^{-1}$ is

(A) 270 (B) 70 (C) 200 (D) 240 mol

(5) The correct formula for hexaaminecobalt (III) nitrate is

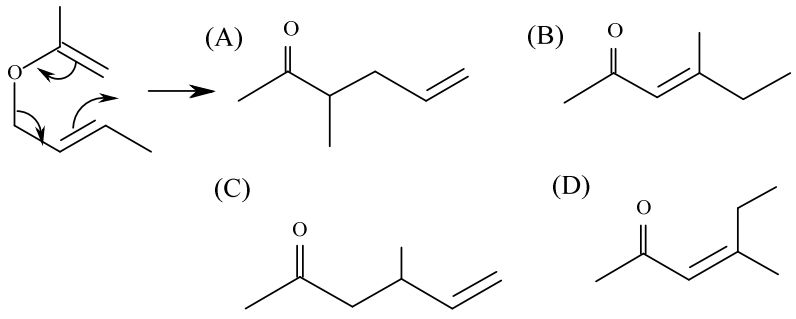
(A) $[Co_3(NH_3)](NO_3)_3$ (B) $[Co_3(NH_3)_6](NO_3)_3$

(C) $[Co(NO_3)_3].6NH_3$ (D) $[Co(NH_3)_6](NO_3)_3$

(6) For the reaction $PCl_3(g) + Cl_2(g)$ Â $PCl_5(g)$, K_c is 26 at 250°C. K_p at the same temperature is ($R = 8.314\ J\ K^{-1}mol^{-1}$)

(A) 4.6×10^{-3} (B) 5.7×10^{-3} (C) 6.0×10^{-3} (D) 8.3×10^{-3}

(7) Curved arrows are used in Organic Chemistry to show the movements of electrons in the mechanism of a reaction. The correct product of the following reaction is



(8) Denaturation of protein due to change in pH could be due to

(A) loss of van der Waal's interaction (B) hydrophobic interaction

(C) change in ionic interaction (D) breaking of covalent bonds

(9) The initial activity of a radionuclide is 9750 counts per min and 975 counts after 5min. The decay constant of the radionuclide in min^{-1} is about

(A) 0.23 (B) 0.46 (C) 0.69 (D) 0.99

(10) According to VSEPR theory the shape of IF_5 molecule will be

(A) tetrahedral (B) trigonal bipyramid (C) square pyramid (D) trigonal planar

(11) The formal charges on the atoms underlined are

C_6H_5-<u>C</u>≡<u>N</u>-<u>O</u>

(A) C = 0, N = –1, O = +1 (B) C = –1, N = +1, O = –1

(C) C = 0, N = +1, O = –1 (D) C = +1, N = 0, O = –1

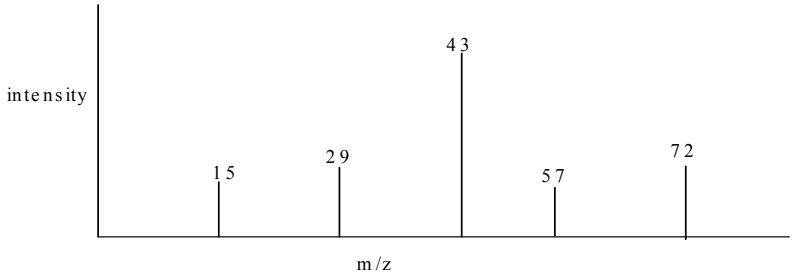
(12) The number of α-particles emitted per second by a radioactive element reduces to 6.25% of the original value in 48 days. The half-life period of the element in days is

(A) 3 (B) 8 (C)12 (D) 16

(13) The compound that **does not** have a π bond is

(A) SO_2 (B) SF_6 (C) O_2 (D) SO_3

(14) In mass spectrometry a compound is bombarded with high energy electrons to break it into smaller fragments, which are recorded in the form of their masses (m/z). For example butane gives fragments like m/z 58, 43, 29, 15, etc. The mass spectrum of an unknown compound is shown below.



The likely compound among the following is

(A) CH_3COCl (B) $CH_2=CH\text{-}CH_2CH_2OH$
(C) CH_3CH_2COOH (D) $CH_3COCH_2CH_3$

(15) The solubility of calcium phosphate is S $moldm^{-3}$. Hence, the solubility product is

(A) S^5 (B) $27S^3$ (C) $54S^4$ (D) $108S^5$

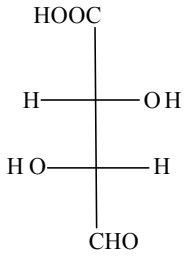
(16) The number of valence electrons in an atom with the configuration $1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^2$ is

(A) 6 (B) 5 (C) 4 (D) 2

(17) The elevation in boling point of a solution containing 13.44g of $CuCl_2$ in 1 kg of water is (K_b = 0.52 K kg mol^{-1})

(A) 0.05 (B) 0.10 (C) 0.16 (D) 0.21

(18) The configurations of the carbon atoms C_2 and C_3 in the following compound are respectively



(A) R, R (B) S, S (C) R, S (D) S, R

(19) 0.1dm^3 of 0.1M acetic acid is titrated against 0.1M NaOH. When 50 cm^3 of 0.1M NaOH are added, the pH of the solution will be (pK_a = 4.74)

(A) 2.37 (B) 4.74 (C) 1.34 (D) 5.74

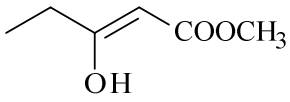
(20) The IUPAC name of complex $[Cu(en)_2(H_2O)_2]^+$ is

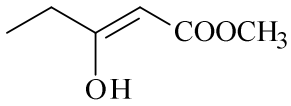
(A) ethylene diamine Cu(II) dihydrate (B) diaquobis(ethylenediamine) Cu(II) ion
(C) diaquobisdiethylamineCu(II) ion (D) diaquobis(ethylenediamine) cuprate (II)

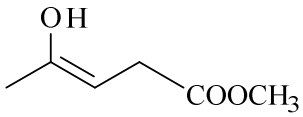
(21) Two protein molecules with the same average molecular mass (molecular weight) can absorb different amount of ultraviolet radiation due to difference in the content of

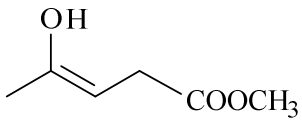
(A) tyrosine (B) glutamic acid (C) lysine (D) methionine

(22) Each of the following options contains a structure and a description indicating the existence of given structure. The correct option for methyl 3-hydroxypent-2-enoate is

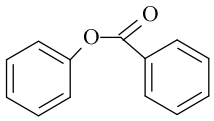
(A)  YES

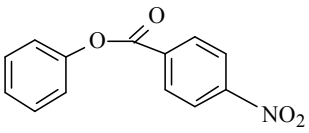
(B)  NO

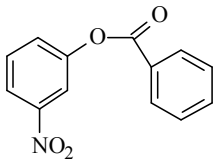
(C)  NO

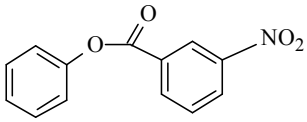
(D)  YES

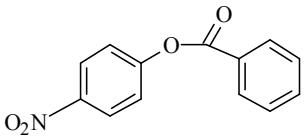
(23) Major product of mononitration of the following compound is



(A) 

(B) 

(C) 

(D) 

(24) If $a \neq b \neq c$ and $\alpha = \beta = \gamma = 90^{o}$, the crystal system is

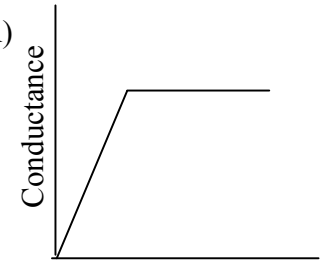
(A) monoclinic (B) triclinic (C) hexagonal (D) orthorhombic

(25) The electronic spectrum of $[Ni(H_2O)_6]^{++}$ shows a band at 8500 cm^{-1} due to d–d transition. $[Ph_4As]_2[NiCl_4]$ will have such a transition in cm^{-1} at

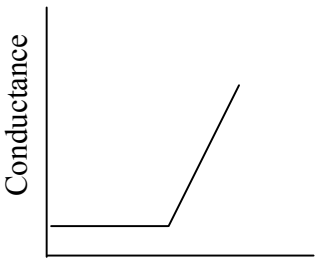
(A) 3778 (B) 8500 (C) 4250 (D) 850

(26) In the conductometric titration of silver nitrate against KCl, the graph obtained is

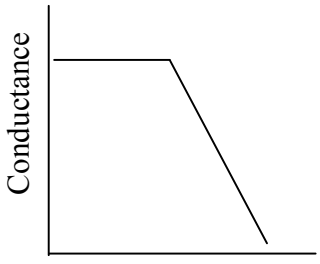
(A)



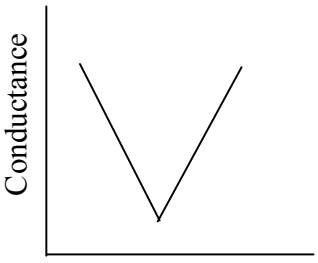
(B)



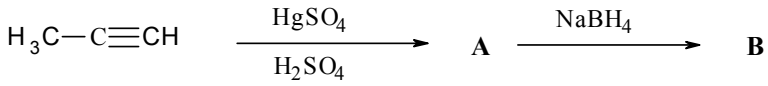
(C)



(D)



(27) The product obtained from the following sequence of reactions is



(A) propanal (B) 2-propanol (C) 1-propanol (D) propane

(28) The compound in which Mn has oxidation state of +3 is

(A) $KMnO_4$ (B) $K_2[Mn(CN)_6]$ (C) $MnSO_4$ (D) $CsMn(SO_4)_2.12H_2O$

(29) The SI units of viscocity is

(A) Nsm^2 (B) Ns^2m (C) Nsm^{-2} (D) $Ns^{-2}m$

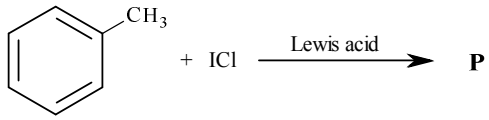
(30) If titration of an amino acid present in the solution yielded pI (isoelectric point) value of 10.80, the amino acid present in the solution may be

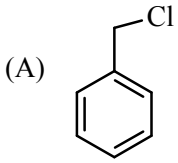
(A) glycine (B) arginine (C) histidine (D) proline

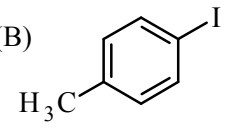
(31) In the coordination compound, $Na_2[Pt(CN)_4]$, the Lewis acid is

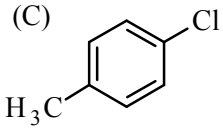
(A) $[Pt(CN)_4]^{2-}$ (B) Na^+ (C) Pt^{2+} (D) CN^-

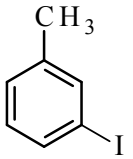
(32) The product (P) of the following reaction is –



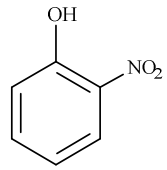
(A) 

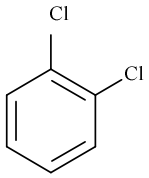
(B) 

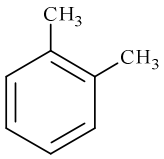
(C) 

(D) 

(33) The correct order of dipole moment for the following molecules is

(I) 

(II) 

(III) 

(A) I = II = III (B) I < II < III (C) I > II > III (D) II < III < I

(34) Lead dissolves most readily in dilute

(A) acetic acid (B) sulphuric acid (C) phosphoric acid (D) sodium hydroxide

(35) The degrees of freedom for the system $CaCO_3(s)$ Â $CaO(s) + CO_2(g)$ are

(A) 1 (B) 2 (C) 3 (D) 4

(36) Semipermeable nature of the cell membrane can be attributed to the presence of

(A) protein and DNA (B) lipid and protein

(C) polysaccharide and lipid (D) DNA and lipid

(37) The emf of the cell $(Zn \mid ZnSO_4\ (0.1M) \parallel CdSO_4\ (0.01M) \mid Cd)$ is

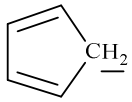
$(E^0_{Zn^{2+}|Zn} = -0.76\ V,\ E^0_{Cd^{2+}|Cd} = 0.40\ V$ at 298 K)

(A) +0.33 V (B) +0.36 V (C) +1.13 V (D) –0.36 V

(38) The nitrogen compound formed when $Ca(CN)_2$ reacts with steam or hot water is

(A) N_2O (B) NO (C) NO_2 (D) NH_3

(39) The order of acidities of the H-atoms underlined in the following compounds is in the order-

(I) Ph- $C\underline{H_2}$-CH_3 (II) Ph-C≡C$\underline{H}$ (III) Ph-CH=C$\underline{H_2}$ (IV) 

(A) IV > II > I > III (B) II > IV > III > I (C) III > IV > I > II (D) I > III > II > IV

(40) The half time for a second order reaction with equal concentrations of the reactants is 35 seconds. 99% reaction will be completed in

(A) 69s (B) 138s (C) 1733s (D) 3465s

(41) The 'd' orbitals will be split under square planar geometry into

(A) two levels (B) three levels (C) four levels (D) five levels

(42) Rotational spectra of molecules are observed in

(A) UV region (B) Visible region

(C) Near infrared region (D) Far infrared region

(43) The pair of cations which **cannot** be separated by H_2S in a 0.3N acid solution is

(A) Al^{+++}, Hg^{++} (B) Bi^{+++}, Pb^{++} (C) Zn^{++}, Cu^{++} (D) Ni^{++}, Cd^{++}

(44) Structural features of proteins secreted outside the cells may be stabilised by presence of

(A) hydrogen bond (B) disulfide bond

(C) hydrophobic force (D) phospho-diester bond

(45) The C-O-C bond angle in dimethyl ether is

(A) $109^o28'$ (B) 110^o (C) 120^o (D) 180^o

(46) Dimethyl glyoxime forms a square planar complex with Ni^{2+}. This complex should be

(A) diamagnetic

(B) paramagnetic having 1 unpaired electron

(C) paramagnetic having 2 unpaired electrons

(D) ferromagnetic

(47) A 0.056 M solution of benzoic acid, C_6H_5COOH, is titrated with a strong base. $[H^+]$ of the solution when half of the solution is titrated before the equivalence point is (Ka of benzoic acid = 6.3 x 10^{-5})

(A) 6.3×10^{-5} M (B) 1.8×10^{-3} M (C) 7.9×10^{-3} M (D) 2.6×10^{-2} M

(48) The formula of the isothiocyanate ion is

(A) OCN^- (B) SCN^- (C) ONC^- (D) CN^-

(49) The compound that is chiral is

(A) 3-methyl-3-hexene (B) 4-chloro-1-methylcyclohexane

(C) 2-phenylpentane (D) 1,3-diisopropylbenzene

(50) The monomer/s of the following polymer is/are

$(-CH_2-CH(CH_3)-CH_2-CH(CH_3)-CH_2-CH(CH_3)-)n$

(A) ethylene (B) propylene (C) 2-butene (D) ethylene + propylene

(51) Of the interhalogen compounds, ClF_3 is more reactive than BrF_3, but BrF_3 has higher conductance in the liquid state. The reason is that

(A) BrF_3 has higher molecular weight (B) ClF_3 is volatile

(C) BrF_3 dissociates into BrF^{2-} and BrF^{4-} more easily (D) ClF_3 is most reactive

(52) An element X is found to combine with oxygen to form X_4O_6. If 8.40 g of this element combine with 6.50 g of oxygen, the atomic weight of the element in grams is

(A) 24.0 (B) 31.0 (C) 50.4 (D) 118.7

(53) Synthesis of RNA in a cell would be affected adversely due to shortage of

(A) sulfate (B) acetate (C) oxalate (D) phosphate

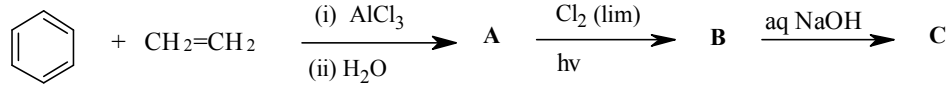
(54) The most abundant element in the earth's crust is

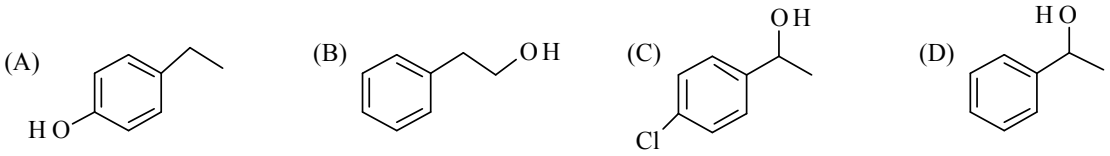
(A) aluminium (B) oxygen (C) silicon (D) iron

(55) A beaker is heated from 27°C to 127°C, the percentage of air originally present in beaker that is expelled is

(A) 50% (B) 25% (C) 33% (D) 40%

(56) The product (C) of the following sequence of reactions is





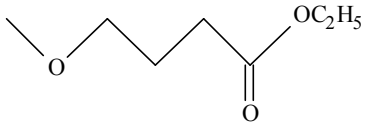
(57) The strongest, but the most reactive bond among the following is –

(A) C=N (B) C=C (C) C≡C (D) C=O

(58) Radioactive inert gas is –

(A) technetium (B) radon (C) xenon (D) curium

(59) The IUPAC name of the following compound is –



(A) 3-methoxy ethylpropanoate (B) ethyl 4-methoxybutanoate

(C) 1,4-diethoxybutane (D) ethoxy 3-methoxybutyrate

(60) Excess of silver nitrate is added to a water sample to determine the amount of chloride ion present in the sample. 1.4 g of silver chloride is precipitated. The mass of chloride ion present in the sample is

Molar masses ($g \cdot mol^{-1}$): $AgNO_3$ 169.91, AgCl 143.25

(A) 0.25 g (B) 0.35 g (C) 0.50 g (D) 0.75 g

(61) The best nucleophile among the following is –

(A) H_2O (B) CH_3SH (C) Cl^- (D) NH_3

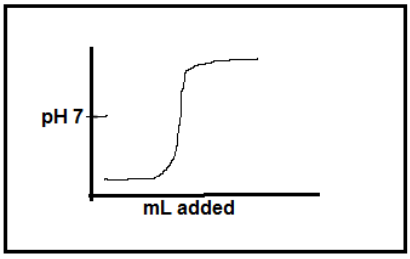
(62) The wavelength of a moving body of mass 0.1mg is 3.31×10^{-29} m. The kinetic energy of the body in J would be

(A) 2.0×10^{-6} (B) 1.0×10^{-3} (C) 4.0×10^{-3} (D) 2.0×10^{-3}

(63) Secondary structures could be formed in nucleic acid similar to protein due to formation of

(A) covalent bond (B) ionic bond (C) co-ordinate bond (D) hydrogen bond

(64) The following titration curve represents the titration of a _____acid with a ____base.



(A) strong, strong (B) weak, strong (C) strong, weak (D) weak, weak

(65) The element with the lowest electronegativity is

(A) S (B) I (C) Ba (D) Al

(66) Oxalic acid, $H_2C_2O_4$, reacts with permanganate ion according to the balanced equation $5H_2C_2O_4(aq) + 2MnO_4^-(aq) + 6H^+(aq)$ Â $2Mn^{2+}(aq) + 10CO_2(g) + 8H_2O(l)$. The volume in mL of 0.0162M $KMnO_4$ solution required to react with 25.0 mL of 0.022 M $H_2C_2O_4$ solution is

(A) 13.6 (B) 18.5 (C) 33.8 (D) 84.4

(67) The element that has the highest tendency to catenate is –

(A) silicon (B) germanium (C) sulphur (D) boron

(68) The isotope of carbon which is used in carbon dating (a method to estimate the age of an ancient sample containing carbon) is-

(A) carbon-12 (B) carbon-13 (C) carbon-14 (D) carbon-15

(69) Electronic configurations for the atoms of four elements are given below. The configuration that indicates colourless aqueous solution is

(A) 2,8,14,2 (B) 2,8,16,2 (C) 2,8,18,2 (D) 2,8,13,1

(70) The number of stereoisomers of compound CH_3-CH=CH-CH(Br)CH_3 is –

(A) 2 (B) 3 (C) 4 (D) 6

(71) At 445° C, Kc for the following reaction is 0.020.

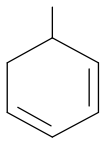
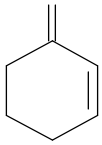
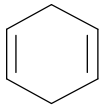
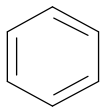
2HI(g) Â $H_2(g)$ + $I_2(g)$

A mixture of H_2, I_2, and HI in a vessel at 445°C has the following concentrations:

[HI] = 2.0M, [H_2] = 0.50M and [I_2] = 0.10 M. The statement that is true concerning the reaction quotient, Q_c is

(A) $Q_c = K_c$; the system is at equilibrium
(B) Q_c is less than K_c; more H_2 and I_2 will be produced
(C) Q_c is less than K_c; more HI will be produced
(D) Q_c is greater than K_c; more H_2 and I_2 will be produced

(72) The order of decreasing stability is

(I)  (II)  (III)  (IV) 

(A) IV>I>II>III (B) I>IV>III>II (C) I>II>IV>III (D) IV>II> I> III

(73) The number of amino acid residues found in a protein that is synthesized from a RNA molecule with 120 nucleotides is

(A) 120 (B) 80 (C) 40 (D) 60

(74) Hypochlorous acid ionizes as

$HOCl(aq)$ Â $H^+(aq) + OCl^-(aq)$.

$OCl^-(aq) + H_2O(l)$ Â $HOCl(aq) + OH^-(aq)$

K_a for this reaction at 25°C is 3.0×10^{-8} ($Kw = 1.0 \times 10^{-14}$ at 25°C)

Hence, K_h for HOCl is

(A) 3.3×10^{-7} (B) 3.0×10^{-8} (C) 3.0×10^{6} (D) 3.3×10^{7}

(75) Einsteinium has 11 electrons in the 4f subshell. The number of unpaired electrons in the subshell is-

(A) 3 (B) 4 (C) 7 (D) 11

(76) The order of reactivity of ammonia with the following compounds is

CH_2=CHBr (II) CH_3-CH_2-COCl (III) CH_3-CH_2-CH_2-Cl (IV) $(CH_3)_3$C-Br

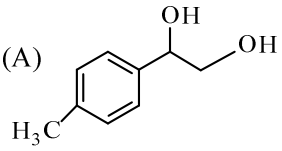
(A) IV>II>I>III (B) II>IV>III>I (C) III>IV>II>I (D) I>IV>II>III

(77) The freezing point of a solution containing 8.1g of HBr in 100 g of water, assuming the acid to be 90% ionized is [H=1, Br=80, K_f for water = 1.86 K kg mol^{-1}]

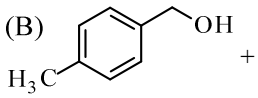
(A) 0.85^0C (B) -3.53^0C (C) 0^0C (D) -0.35^0C

(78) The reaction of 50% aq KOH on an equimolar mixture of 4-methylbenzaldehyde and formaldehyde followed by acidification gives-

(A)

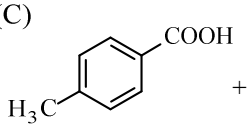


(B)



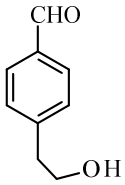
+ HCOOH

(C)



+ CH_3OH

(D)



(79) Iodide ion is oxidized by acidified dichromate ions as shown in this equation.

$Cr_2O_7^{2-}(aq) + 9I^-(aq) + 14H^+(aq) \rightarrow 2Cr^{3+}(aq) + 3I_3^-(aq) + 7H_2O(l)$. These data were obtained when the reaction was studied at a constant pH. The order of the reaction with respect to $Cr_2O_7^{2-}(aq)$ and $I^-(aq)$ are

Experiment	$[Cr_2O_7^{2-}]$,M	$[I^-]$, M	Rate,M.s $^{-1}$
1	0.0050	0.0125	0.00050
2	0.010	0.0125	0.0010
3	0.0150	0.0250	0.0060

(A) first order with respect to both $Cr_2O_7^{2-}$ and I^-

(B) second order with respect to both $Cr_2O_7^{2-}$ and I^-

(C) second order with respect to $Cr_2O_7^{2-}$ and first order with respect to I^-

(D) first order with respect to $Cr_2O_7^{2-}$ and second order with respect to I^-

(80) The number of atoms per unit cell and the number of the nearest neighbours in a body centred cubic structure are-

(A) 4, 12 (B) 2,6 (C) 9, 6 (D) 2, 8

Periodic Table of Elements

1																	18
1 1.00794 H Hydrogen	2											13	14	15	16	17	2 4.002602 He Helium
3 6.941 Li Lithium	4 9.012082 Be Beryllium											5 10.811 B Boron	6 12.011 C Carbon	7 14.007 N Nitrogen	8 15.999 O Oxygen	9 18.998 F Fluorine	10 20.180 Ne Neon
11 22.990 Na Sodium	12 24.305 Mg Magnesium	3	4	5	6	7	8	9	10	11	12	13 26.982 Al Aluminium (Aluminum)	14 28.086 Si Silicon	15 30.974 P Phosphorus	16 32.066 S Sulphur (Sulfur)	17 35.453 Cl Chlorine	18 39.948 Ar Argon
19 39.098 K Potassium (Kalium)	20 40.078 Ca Calcium	21 44.956 Sc Scandium	22 47.87 Ti Titanium	23 50.942 V Vanadium	24 51.996 Cr Chromium	25 54.938 Mn Manganese	26 55.845 Fe Iron (Ferrum)	27 58.933 Co Cobalt	28 58.693 Ni Nickel	29 63.546 Cu Copper (Cuprum)	30 65.39 Zn Zinc	31 69.723 Ga Gallium	32 72.61 Ge Germanium	33 74.922 As Arsenic	34 78.96 Se Selenium	35 79.904 Br Bromine	36 83.80 Kr Krypton
37 85.468 Rb Rubidium	38 87.62 Sr Strontium	39 88.906 Y Yttrium	40 91.224 Zr Zirconium	41 92.906 Nb Niobium	42 95.94 Mo Molybdenum	43 [97.907] Tc Technetium	44 101.07 Ru Ruthenium	45 102.91 Rh Rhodium	46 106.42 Pd Palladium	47 107.87 Ag Silver (Argentum)	48 112.41 Cd Cadmium	49 114.82 In Indium	50 118.71 Sn Tin (Stannum)	51 121.76 Sb Antimony (Stibium)	52 127.60 Te Tellurium	53 126.90 I Iodine	54 131.29 Xe Xenon
55 132.91 Cs Caesium	56 137.33 Ba Barium	57 138.91 La Lanthanum	72 178.49 Hf Hafnium	73 180.95 Ta Tantalum	74 183.84 W Tungsten (Wolfram)	75 186.21 Re Rhenium	76 190.23 Os Osmium	77 192.22 Ir Iridium	78 195.08 Pt Platinum	79 196.97 Au Gold (Aurum)	80 200.59 Hg Mercury (Hydrargyrum)	81 204.38 Tl Thallium	82 207.20 Pb Lead (Plumbum)	83 208.98 Bi Bismuth	84 [208.98] Po Polonium	85 [209.99] At Astatine	86 [222.02] Rn Radon
87 [223.02] Fr Francium	88 [226.03] Ra Radium	89 [227.03] Ac Actinium	104 [261.11] Rf Rutherfordium	105 [262.11] Db Dubnium	106 [263.12] Sg Seaborgium	107 [262.12] Bh Bohrium	108 [265.13] Hs Hassium	109 [266] Mt Meitnerium	110 [269] Uun Ununnilium	111 [272] Uuu Ununonium	112 [277] Uub Ununbium		114 [285] Uuq Ununquadium		116 [289] Uuh Ununhexium		118 [293] Uuo Ununoctium

Transition Elements

Alkali Metals

Alkaline Earths

Halogens

Noble Gases

Lanthanides (Rare Earths)														
58 140.12 Ce Cerium	59 140.91 Pr Praseodymium	60 144.24 Nd Neodymium	61 [144.91] Pm Promethium	62 150.36 Sm Samarium	63 151.96 Eu Europium	64 157.25 Gd Gadolinium	65 158.93 Tb Terbium	66 162.50 Dy Dysprosium	67 164.93 Ho Holmium	68 167.26 Er Erbium	69 168.93 Tm Thulium	70 173.04 Yb Ytterbium	71 174.97 Lu Lutetium
90 232.04 Th Thorium	91 231.04 Pa Protactinium	92 238.03 U Uranium	93 [237.05] Np Neptunium	94 [244.06] Pu Plutonium	95 [243.06] Am Americium	96 [247] Cm Curium	97 [247.07] Bk Berkelium	98 [251.08] Cf Californium	99 [252.08] Es Einsteinium	100 [257.10] Fm Fermium	101 [258.10] Md Mendelevium	102 [259.10] No Nobelium	103 [262.11] Lr Lawrencium

Actinides

Key

atomic no. mass [mass: longest lived isotope]

Element Symbol

name (name: historical or alternative name)

Logarithms

	0	1	2	3	4	5	6	7	8	9	Mean Differences								
											1	2	3	4	5	6	7	8	9
10	0000	0043	0086	0128	0170	0212	0253	0294	0334	0374	4	8	12	17	21	25	29	33	37
11	0414	0453	0492	0531	0569	0607	0645	0682	0719	0755	4	8	11	15	19	23	26	30	34
12	0792	0828	0864	0899	0934	0969	1004	1038	1072	1106	3	7	10	14	17	21	24	28	31
13	1139	1173	1206	1239	1271	1303	1335	1367	1399	1430	3	6	10	13	16	19	23	26	29
14	1461	1492	1523	1553	1584	1614	1644	1673	1703	1732	3	6	9	12	15	18	21	24	27
15	1761	1790	1818	1847	1875	1903	1931	1959	1987	2014	3	6	8	11	14	17	20	22	25
16	2041	2068	2095	2122	2148	2175	2201	2227	2253	2279	3	5	8	11	13	16	18	21	24
17	2304	2330	2355	2380	2405	2430	2455	2480	2504	2529	2	5	7	10	12	15	17	20	22
18	2553	2577	2601	2625	2648	2672	2695	2718	2742	2765	2	5	7	9	12	14	16	19	21
19	2788	2810	2833	2856	2878	2900	2923	2945	2967	2989	2	4	7	9	11	13	16	18	20
20	3010	3032	3054	3075	3096	3118	3139	3160	3181	3201	2	4	6	8	11	13	15	17	19
21	3222	3243	3263	3284	3304	3324	3345	3365	3385	3404	2	4	6	8	10	12	14	16	18
22	3424	3444	3464	3483	3502	3522	3541	3560	3579	3598	2	4	6	8	10	12	14	15	17
23	3617	3636	3655	3674	3692	3711	3729	3747	3766	3784	2	4	6	7	9	11	13	15	17
24	3802	3820	3838	3856	3874	3892	3909	3927	3945	3962	2	4	5	7	9	11	12	14	16
25	3979	3997	4014	4031	4048	4065	4082	4099	4116	4133	2	3	5	7	9	10	12	14	15
26	4150	4166	4183	4200	4216	4232	4249	4265	4281	4298	2	3	5	7	8	10	11	13	15
27	4314	4330	4346	4362	4378	4393	4409	4425	4440	4456	2	3	5	6	8	9	11	13	14
28	4472	4487	4502	4518	4533	4548	4564	4579	4594	4609	2	3	5	6	8	9	11	12	14
29	4624	4639	4654	4669	4683	4698	4713	4728	4742	4757	1	3	4	6	7	9	10	12	13
30	4771	4786	4800	4814	4829	4843	4857	4871	4886	4900	1	3	4	6	7	9	10	11	13
31	4914	4928	4942	4955	4969	4983	4997	5011	5024	5038	1	3	4	6	7	8	10	11	12
32	5051	5065	5079	5092	5105	5119	5132	5145	5159	5172	1	3	4	5	7	8	9	11	12
33	5185	5198	5211	5224	5237	5250	5263	5276	5289	5302	1	3	4	5	6	8	9	10	12
34	5315	5328	5340	5353	5366	5378	5391	5403	5416	5428	1	3	4	5	6	8	9	10	11
35	5441	5453	5465	5478	5490	5502	5514	5527	5539	5551	1	2	4	5	6	7	9	10	11
36	5563	5575	5587	5599	5611	5623	5635	5647	5658	5670	1	2	4	5	6	7	8	10	11
37	5682	5694	5705	5717	5729	5740	5752	5763	5775	5786	1	2	3	5	6	7	8	9	10
38	5798	5809	5821	5832	5843	5855	5866	5877	5888	5899	1	2	3	5	6	7	8	9	10
39	5911	5922	5933	5944	5955	5966	5977	5988	5999	6010	1	2	3	4	5	7	8	9	10
40	6021	6031	6042	6053	6064	6075	6085	6096	6107	6117	1	2	3	4	5	6	8	9	10
41	6128	6138	6149	6160	6170	6180	6191	6201	6212	6222	1	2	3	4	5	6	7	8	9
42	6232	6243	6253	6263	6274	6284	6294	6304	6314	6325	1	2	3	4	5	6	7	8	9
43	6335	6345	6355	6365	6375	6385	6395	6405	6415	6425	1	2	3	4	5	6	7	8	9
44	6435	6444	6454	6464	6474	6484	6493	6503	6513	6522	1	2	3	4	5	6	7	8	9
45	6532	6542	6551	6561	6571	6580	6590	6599	6609	6618	1	2	3	4	5	6	7	8	9
46	6628	6637	6646	6656	6665	6675	6684	6693	6702	6712	1	2	3	4	5	6	7	7	8
47	6721	6730	6739	6749	6758	6767	6776	6785	6794	6803	1	2	3	4	5	5	6	7	8
48	6812	6821	6830	6839	6848	6857	6866	6875	6884	6893	1	2	3	4	4	5	6	7	8
49	6902	6911	6920	6928	6937	6946	6955	6964	6972	6981	1	2	3	4	4	5	6	7	8
50	6990	6998	7007	7016	7024	7033	7042	7050	7059	7067	1	2	3	3	4	5	6	7	8
51	7076	7084	7093	7101	7110	7118	7126	7135	7143	7152	1	2	3	3	4	5	6	7	8
52	7160	7168	7177	7185	7193	7202	7210	7218	7226	7235	1	2	2	3	4	5	6	7	7
53	7243	7251	7259	7267	7275	7284	7292	7300	7308	7316	1	2	2	3	4	5	6	6	7
54	7324	7332	7340	7348	7356	7364	7372	7380	7388	7396	1	2	2	3	4	5	6	6	7

No.	log		No.	log
$\pi = 3{\cdot}14159$	0·49715	$\ln x = \log_e x = (1/M)\log_{10} x$	$(1/M) = 2{\cdot}30259$	0·36222
$e = 2{\cdot}71828$	0·43429	$\log x = \log_{10} x = M \log_e x$	$M = 0{\cdot}43429$	$\bar{1}{\cdot}63778$

Logarithms

	0	1	2	3	4	5	6	7	8	9	Mean Differences								
											1	2	3	4	5	6	7	8	9
55	7404	7412	7419	7427	7435	7443	7451	7459	7466	7474	1	2	2	3	4	5	5	6	7
56	7482	7490	7497	7505	7513	7520	7528	7536	7543	7551	1	2	2	3	4	5	5	6	7
57	7559	7566	7574	7582	7589	7597	7604	7612	7619	7627	1	2	2	3	4	5	5	6	7
58	7634	7642	7649	7657	7664	7672	7679	7686	7694	7701	1	1	2	3	4	4	5	6	7
59	7709	7716	7723	7731	7738	7745	7752	7760	7767	7774	1	1	2	3	4	4	5	6	7
60	7782	7789	7796	7803	7810	7818	7825	7832	7839	7846	1	1	2	3	4	4	5	6	6
61	7853	7860	7868	7875	7882	7889	7896	7903	7910	7917	1	1	2	3	4	4	5	6	6
62	7924	7931	7938	7945	7952	7959	7966	7973	7980	7987	1	1	2	3	3	4	5	6	6
63	7993	8000	8007	8014	8021	8028	8035	8041	8048	8055	1	1	2	3	3	4	5	5	6
64	8062	8069	8075	8082	8089	8096	8102	8109	8116	8122	1	1	2	3	3	4	5	5	6
65	8129	8136	8142	8149	8156	8162	8169	8176	8182	8189	1	1	2	3	3	4	5	5	6
66	8195	8202	8209	8215	8222	8228	8235	8241	8248	8254	1	1	2	3	3	4	5	5	6
67	8261	8267	8274	8280	8287	8293	8299	8306	8312	8319	1	1	2	3	3	4	5	5	6
68	8325	8331	8338	8344	8351	8357	8363	8370	8376	8382	1	1	2	3	3	4	4	5	6
69	8388	8395	8401	8407	8414	8420	8426	8432	8439	8445	1	1	2	2	3	4	4	5	6
70	8451	8457	8463	8470	8476	8482	8488	8494	8500	8506	1	1	2	2	3	4	4	5	6
71	8513	8519	8525	8531	8537	8543	8549	8555	8561	8567	1	1	2	2	3	4	4	5	5
72	8573	8579	8585	8591	8597	8603	8609	8615	8621	8627	1	1	2	2	3	4	4	5	5
73	8633	8639	8645	8651	8657	8663	8669	8675	8681	8686	1	1	2	2	3	4	4	5	5
74	8692	8698	8704	8710	8716	8722	8727	8733	8739	8745	1	1	2	2	3	4	4	5	5
75	8751	8756	8762	8768	8774	8779	8785	8791	8797	8802	1	1	2	2	3	3	4	5	5
76	8808	8814	8820	8825	8831	8837	8842	8848	8854	8859	1	1	2	2	3	3	4	5	5
77	8865	8871	8876	8882	8887	8893	8899	8904	8910	8915	1	1	2	2	3	3	4	4	5
78	8921	8927	8932	8938	8943	8949	8954	8960	8965	8971	1	1	2	2	3	3	4	4	5
79	8976	8982	8987	8993	8998	9004	9009	9015	9020	9025	1	1	2	2	3	3	4	4	5
80	9031	9036	9042	9047	9053	9058	9063	9069	9074	9079	1	1	2	2	3	3	4	4	5
81	9085	9090	9096	9101	9106	9112	9117	9122	9128	9133	1	1	2	2	3	3	4	4	5
82	9138	9143	9149	9154	9159	9165	9170	9175	9180	9186	1	1	2	2	3	3	4	4	5
83	9191	9196	9201	9206	9212	9217	9222	9227	9232	9238	1	1	2	2	3	3	4	4	5
84	9243	9248	9253	9258	9263	9269	9274	9279	9284	9289	1	1	2	2	3	3	4	4	5
85	9294	9299	9304	9309	9315	9320	9325	9330	9335	9340	1	1	2	2	3	3	4	4	5
86	9345	9350	9355	9360	9365	9370	9375	9380	9385	9390	1	1	2	2	3	3	4	4	5
87	9395	9400	9405	9410	9415	9420	9425	9430	9435	9440	0	1	1	2	2	3	3	4	4
88	9445	9450	9455	9460	9465	9469	9474	9479	9484	9489	0	1	1	2	2	3	3	4	4
89	9494	9499	9504	9509	9513	9518	9523	9528	9533	9538	0	1	1	2	2	3	3	4	4
90	9542	9547	9552	9557	9562	9566	9571	9576	9581	9586	0	1	1	2	2	3	3	4	4
91	9590	9595	9600	9605	9609	9614	9619	9624	9628	9633	0	1	1	2	2	3	3	4	4
92	9638	9643	9647	9652	9657	9661	9666	9671	9675	9680	0	1	1	2	2	3	3	4	4
93	9685	9689	9694	9699	9703	9708	9713	9717	9722	9727	0	1	1	2	2	3	3	4	4
94	9731	9736	9741	9745	9750	9754	9759	9763	9768	9773	0	1	1	2	2	3	3	4	4
95	9777	9782	9786	9791	9795	9800	9805	9809	9814	9818	0	1	1	2	2	3	3	4	4
96	9823	9827	9832	9836	9841	9845	9850	9854	9859	9863	0	1	1	2	2	3	3	4	4
97	9868	9872	9877	9881	9886	9890	9894	9899	9903	9908	0	1	1	2	2	3	3	4	4
98	9912	9917	9921	9926	9930	9934	9939	9943	9948	9952	0	1	1	2	2	3	3	4	4
99	9956	9961	9965	9969	9974	9978	9983	9987	9991	9996	0	1	1	2	2	3	3	3	4

p	1	2	3	4	5	6	7	8	9	10
$\log e^p$	0·4343	0·8686	$\bar{1}{\cdot}3029$	1·7372	2·1715	2·6058	$\bar{3}{\cdot}0401$	$\bar{3}{\cdot}4744$	$\bar{3}{\cdot}9087$	$\bar{4}{\cdot}3429$
$\log e^{-p}$	$\bar{1}{\cdot}5657$	$\bar{1}{\cdot}1314$	$\bar{2}{\cdot}6971$	$\bar{2}{\cdot}2628$	$\bar{3}{\cdot}8285$	$\bar{3}{\cdot}3942$	$\bar{4}{\cdot}9599$	$\bar{4}{\cdot}5256$	$\bar{4}{\cdot}0913$	$\bar{5}{\cdot}6571$

Antilogarithms

	0	1	2	3	4	5	6	7	8	9	Mean Differences								
											1	2	3	4	5	6	7	8	9
·00	1000	1002	1005	1007	1009	1012	1014	1016	1019	1021	0	0	1	1	1	1	2	2	2
·01	1023	1026	1028	1030	1033	1035	1038	1040	1042	1045	0	0	1	1	1	1	2	2	2
·02	1047	1050	1052	1054	1057	1059	1062	1064	1067	1069	0	0	1	1	1	1	2	2	2
·03	1072	1074	1076	1079	1081	1084	1086	1089	1091	1094	0	0	1	1	1	1	2	2	2
·04	1096	1099	1102	1104	1107	1109	1112	1114	1117	1119	0	1	1	1	1	2	2	2	2
·05	1122	1125	1127	1130	1132	1135	1138	1140	1143	1146	0	1	1	1	1	2	2	2	2
·06	1148	1151	1153	1156	1159	1161	1164	1167	1169	1172	0	1	1	1	1	2	2	2	2
·07	1175	1178	1180	1183	1186	1189	1191	1194	1197	1199	0	1	1	1	1	2	2	2	2
·08	1202	1205	1208	1211	1213	1216	1219	1222	1225	1227	0	1	1	1	1	2	2	2	3
·09	1230	1233	1236	1239	1242	1245	1247	1250	1253	1256	0	1	1	1	1	2	2	2	3
·10	1259	1262	1265	1268	1271	1274	1276	1279	1282	1285	0	1	1	1	1	2	2	2	3
·11	1288	1291	1294	1297	1300	1303	1306	1309	1312	1315	0	1	1	1	2	2	2	2	3
·12	1318	1321	1324	1327	1330	1334	1337	1340	1343	1346	0	1	1	1	2	2	2	2	3
·13	1349	1352	1355	1358	1361	1365	1368	1371	1374	1377	0	1	1	1	2	2	2	3	3
·14	1380	1384	1387	1390	1393	1396	1400	1403	1406	1409	0	1	1	1	2	2	2	3	3
·15	1413	1416	1419	1422	1426	1429	1432	1435	1439	1442	0	1	1	1	2	2	2	3	3
·16	1445	1449	1452	1455	1459	1462	1466	1469	1472	1476	0	1	1	1	2	2	2	3	3
·17	1479	1483	1486	1489	1493	1496	1500	1503	1507	1510	0	1	1	1	2	2	2	3	3
·18	1514	1517	1521	1524	1528	1531	1535	1538	1542	1545	0	1	1	1	2	2	2	3	3
·19	1549	1552	1556	1560	1563	1567	1570	1574	1578	1581	0	1	1	1	2	2	3	3	3
·20	1585	1589	1592	1596	1600	1603	1607	1611	1614	1618	0	1	1	1	2	2	3	3	3
·21	1622	1626	1629	1633	1637	1641	1644	1648	1652	1656	0	1	1	2	2	2	3	3	3
·22	1660	1663	1667	1671	1675	1679	1683	1687	1690	1694	0	1	1	2	2	2	3	3	3
·23	1698	1702	1706	1710	1714	1718	1722	1726	1730	1734	0	1	1	2	2	2	3	3	4
·24	1738	1742	1746	1750	1754	1758	1762	1766	1770	1774	0	1	1	2	2	2	3	3	4
·25	1778	1782	1786	1791	1795	1799	1803	1807	1811	1816	0	1	1	2	2	2	3	3	4
·26	1820	1824	1828	1832	1837	1841	1845	1849	1854	1858	0	1	1	2	2	3	3	3	4
·27	1862	1866	1871	1875	1879	1884	1888	1892	1897	1901	0	1	1	2	2	3	3	3	4
·28	1905	1910	1914	1919	1923	1928	1932	1936	1941	1945	0	1	1	2	2	3	3	4	4
·29	1950	1954	1959	1963	1968	1972	1977	1982	1986	1991	0	1	1	2	2	3	3	4	4
·30	1995	2000	2004	2009	2014	2018	2023	2028	2032	2037	0	1	1	2	2	3	3	4	4
·31	2042	2046	2051	2056	2061	2065	2070	2075	2080	2084	0	1	1	2	2	3	3	4	4
·32	2089	2094	2099	2104	2109	2113	2118	2123	2128	2133	0	1	1	2	2	3	3	4	4
·33	2138	2143	2148	2153	2158	2163	2168	2173	2178	2183	0	1	1	2	2	3	3	4	4
·34	2188	2193	2198	2203	2208	2213	2218	2223	2228	2234	1	1	2	2	3	3	4	4	5
·35	2239	2244	2249	2254	2259	2265	2270	2275	2280	2286	1	1	2	2	3	3	4	4	5
·36	2291	2296	2301	2307	2312	2317	2323	2328	2333	2339	1	1	2	2	3	3	4	4	5
·37	2344	2350	2355	2360	2366	2371	2377	2382	2388	2393	1	1	2	2	3	3	4	4	5
·38	2399	2404	2410	2415	2421	2427	2432	2438	2443	2449	1	1	2	2	3	3	4	4	5
·39	2455	2460	2466	2472	2477	2483	2489	2495	2500	2506	1	1	2	2	3	3	4	5	5
·40	2512	2518	2523	2529	2535	2541	2547	2553	2559	2564	1	1	2	2	3	4	4	5	5
·41	2570	2576	2582	2588	2594	2600	2606	2612	2618	2624	1	1	2	2	3	4	4	5	5
·42	2630	2636	2642	2649	2655	2661	2667	2673	2679	2685	1	1	2	2	3	4	4	5	6
·43	2692	2698	2704	2710	2716	2723	2729	2735	2742	2748	1	1	2	3	3	4	4	5	6
·44	2754	2761	2767	2773	2780	2786	2793	2799	2805	2812	1	1	2	3	3	4	4	5	6
·45	2818	2825	2831	2838	2844	2851	2858	2864	2871	2877	1	1	2	3	3	4	5	5	6
·46	2884	2891	2897	2904	2911	2917	2924	2931	2938	2944	1	1	2	3	3	4	5	5	6
·47	2951	2958	2965	2972	2979	2985	2992	2999	3006	3013	1	1	2	3	3	4	5	5	6
·48	3020	3027	3034	3041	3048	3055	3062	3069	3076	3083	1	1	2	3	4	4	5	6	6
·49	3090	3097	3105	3112	3119	3126	3133	3141	3148	3155	1	1	2	3	4	4	5	6	6

Antilogarithms

Degrees	0	1	2	3	4	5	6	7	8	9	Mean Differences								
											1	2	3	4	5	6	7	8	9
·50	3162	3170	3177	3184	3192	3199	3206	3214	3221	3228	1	1	2	3	4	4	5	6	7
·51	3236	3243	3251	3258	3266	3273	3281	3289	3296	3304	1	2	2	3	4	5	5	6	7
·52	3311	3319	3327	3334	3342	3350	3357	3365	3373	3381	1	2	2	3	4	5	5	6	7
·53	3388	3396	3404	3412	3420	3428	3436	3443	3451	3459	1	2	2	3	4	5	6	6	7
·54	3467	3475	3483	3491	3499	3508	3516	3524	3532	3540	1	2	2	3	4	5	6	6	7
·55	3548	3556	3565	3573	3581	3589	3597	3606	3614	3622	1	2	2	3	4	5	6	7	7
·56	3631	3639	3648	3656	3664	3673	3681	3690	3698	3707	1	2	3	3	4	5	6	7	8
·57	3715	3724	3733	3741	3750	3758	3767	3776	3784	3793	1	2	3	3	4	5	6	7	8
·58	3802	3811	3819	3828	3837	3846	3855	3864	3873	3882	1	2	3	4	4	5	6	7	8
·59	3890	3899	3908	3917	3926	3936	3945	3954	3963	3972	1	2	3	4	5	5	6	7	8
·60	3981	3990	3999	4009	4018	4027	4036	4046	4055	4064	1	2	3	4	5	6	6	7	8
·61	4074	4083	4093	4102	4111	4121	4130	4140	4150	4159	1	2	3	4	5	6	7	8	9
·62	4169	4178	4188	4198	4207	4217	4227	4236	4246	4256	1	2	3	4	5	6	7	8	9
·63	4266	4276	4285	4295	4305	4315	4325	4335	4345	4355	1	2	3	4	5	6	7	8	9
·64	4365	4375	4385	4395	4406	4416	4426	4436	4446	4457	1	2	3	4	5	6	7	8	9
·65	4467	4477	4487	4498	4508	4519	4529	4539	4550	4560	1	2	3	4	5	6	7	8	9
·66	4571	4581	4592	4603	4613	4624	4634	4645	4656	4667	1	2	3	4	5	6	7	9	10
·67	4677	4688	4699	4710	4721	4732	4742	4753	4764	4775	1	2	3	4	5	7	8	9	10
·68	4786	4797	4808	4819	4831	4842	4853	4864	4875	4887	1	2	3	4	6	7	8	9	10
·69	4898	4909	4920	4932	4943	4955	4966	4977	4989	5000	1	2	3	5	6	7	8	9	10
·70	5012	5023	5035	5047	5058	5070	5082	5093	5105	5117	1	2	4	5	6	7	8	9	11
·71	5129	5140	5152	5164	5176	5188	5200	5212	5224	5236	1	2	4	5	6	7	8	10	11
·72	5248	5260	5272	5284	5297	5309	5321	5333	5346	5358	1	2	4	5	6	7	9	10	11
·73	5370	5383	5395	5408	5420	5433	5445	5458	5470	5483	1	3	4	5	6	8	9	10	11
·74	5495	5508	5521	5534	5546	5559	5572	5585	5598	5610	1	3	4	5	6	8	9	10	12
·75	5623	5636	5649	5662	5675	5689	5702	5715	5728	5741	1	3	4	5	7	8	9	10	12
·76	5754	5768	5781	5794	5808	5821	5834	5848	5861	5875	1	3	4	5	7	8	9	11	12
·77	5888	5902	5916	5929	5943	5957	5970	5984	5998	6012	1	3	4	5	7	8	10	11	12
·78	6026	6039	6053	6067	6081	6095	6109	6124	6138	6152	1	3	4	6	7	8	10	11	13
·79	6166	6180	6194	6209	6223	6237	6252	6266	6281	6295	1	3	4	6	7	9	10	11	13
·80	6310	6324	6339	6353	6368	6383	6397	6412	6427	6442	1	3	4	6	7	9	10	12	13
·81	6457	6471	6486	6501	6516	6531	6546	6561	6577	6592	2	3	5	6	8	9	11	12	14
·82	6607	6622	6637	6653	6668	6683	6699	6714	6730	6745	2	3	5	6	8	9	11	12	14
·83	6761	6776	6792	6808	6823	6839	6855	6871	6887	6902	2	3	5	6	8	9	11	13	14
·84	6918	6934	6950	6966	6982	6998	7015	7031	7047	7063	2	3	5	6	8	10	11	13	15
·85	7079	7096	7112	7129	7145	7161	7178	7194	7211	7228	2	3	5	7	8	10	12	13	15
·86	7244	7261	7278	7295	7311	7328	7345	7362	7379	7396	2	3	5	7	8	10	12	13	15
·87	7413	7430	7447	7464	7482	7499	7516	7534	7551	7568	2	3	5	7	9	10	12	14	16
·88	7586	7603	7621	7638	7656	7674	7691	7709	7727	7745	2	4	5	7	9	11	12	14	16
·89	7762	7780	7798	7816	7834	7852	7870	7889	7907	7925	2	4	5	7	9	11	13	14	16
·90	7943	7962	7980	7998	8017	8035	8054	8072	8091	8110	2	4	6	7	9	11	13	15	17
·91	8128	8147	8166	8185	8204	8222	8241	8260	8279	8299	2	4	6	8	9	11	13	15	17
·92	8318	8337	8356	8375	8395	8414	8433	8453	8472	8492	2	4	6	8	10	12	14	15	17
·93	8511	8531	8551	8570	8590	8610	8630	8650	8670	8690	2	4	6	8	10	12	14	16	18
·94	8710	8730	8750	8770	8790	8810	8831	8851	8872	8892	2	4	6	8	10	12	14	16	18
·95	8913	8933	8954	8974	8995	9016	9036	9057	9078	9099	2	4	6	8	10	12	15	17	19
·96	9120	9141	9162	9183	9204	9226	9247	9268	9290	9311	2	4	6	8	11	13	15	17	19
·97	9333	9354	9376	9397	9419	9441	9462	9484	9506	9528	2	4	7	9	11	13	15	17	20
·98	9550	9572	9594	9616	9638	9661	9683	9705	9727	9750	2	4	7	9	11	13	16	18	20
·99	9772	9795	9817	9840	9863	9886	9908	9931	9954	9977	2	5	7	9	11	14	16	18	20

ROUGH SHEET

ROUGH SHEET